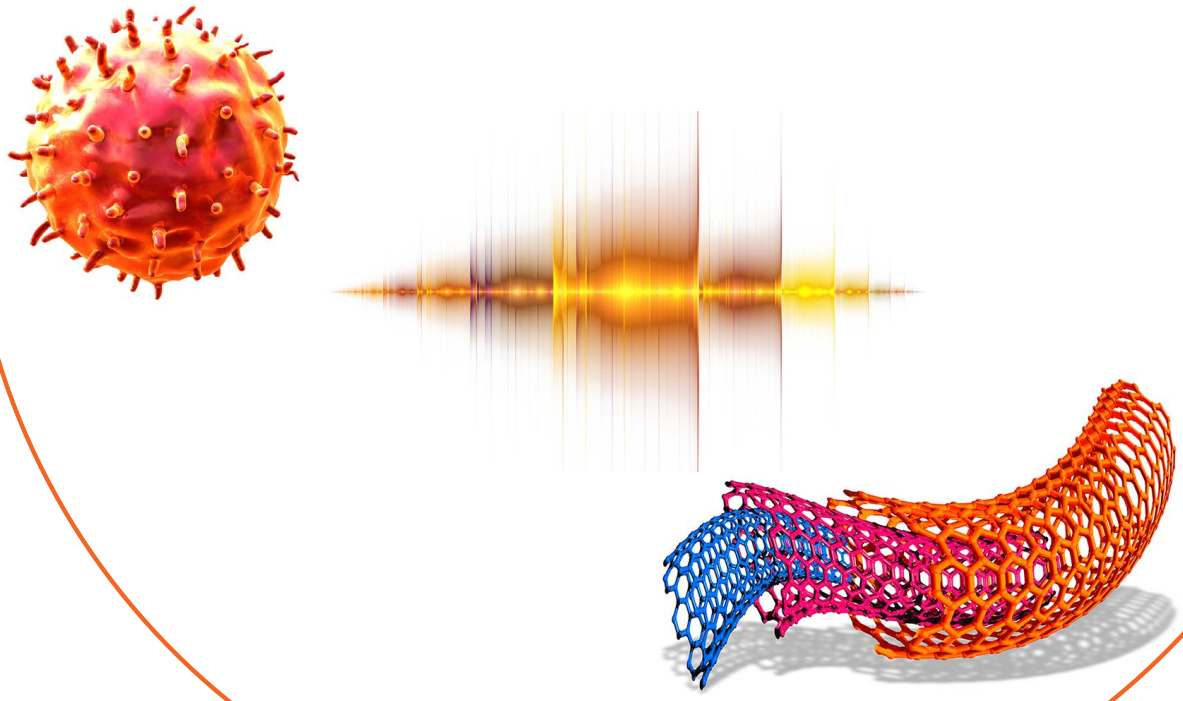
Cambridge Assessment International Education

Syllabus

Cambridge IGCSE™

Combined Science 0653

Use this syllabus for exams in 2023 and 2024.
Exams are available in the June and November series.
Exams are also available in the March series in India only.



Version 1

Cambridge Pathway

Why choose Cambridge International?

Cambridge International prepares school students for life, helping them develop an informed curiosity and a lasting passion for learning. We are part of the University of Cambridge.

Our Cambridge Pathway gives students a clear path for educational success from age 5 to 19. Schools can shape the curriculum around how they want students to learn – with a wide range of subjects and flexible ways to offer them. It helps students discover new abilities and a wider world, and gives them the skills they need for life, so they can achieve at school, university and work.

Our programmes and qualifications set the global standard for international education. They are created by subject experts, rooted in academic rigour and reflect the latest educational research. They provide a strong platform for learners to progress from one stage to the next, and are well supported by teaching and learning resources.

Our mission is to provide educational benefit through provision of international programmes and qualifications for school education and to be the world leader in this field. Together with schools, we develop Cambridge learners who are confident, responsible, reflective, innovative and engaged – equipped for success in the modern world.

Every year, nearly a million Cambridge students from 10 000 schools in 160 countries prepare for their future with the Cambridge Pathway.

'We think the Cambridge curriculum is superb preparation for university.'

Christoph Guttentag, Dean of Undergraduate Admissions, Duke University, USA

Quality management

Cambridge International is committed to providing exceptional quality. In line with this commitment, our quality management system for the provision of international qualifications and education programmes for students aged 5 to 19 is independently certified as meeting the internationally recognised standard, ISO 9001:2015. Learn more at **www.cambridgeinternational.org/ISO9001**

Contents

Important: Changes to this syllabus

For information about changes to this syllabus for 2023 and 2024, go to page 62.

The latest syllabus is version 1, published September 2020. There are no significant changes which affect teaching.

Any textbooks endorsed to support the syllabus for examination from 2019 are still suitable for use with this syllabus.

1 Why choose this syllabus?

Key benefits

Cambridge IGCSE is the world's most popular international qualification for 14 to 16 year olds, although it can be taken by students of other ages. It is tried, tested and trusted.

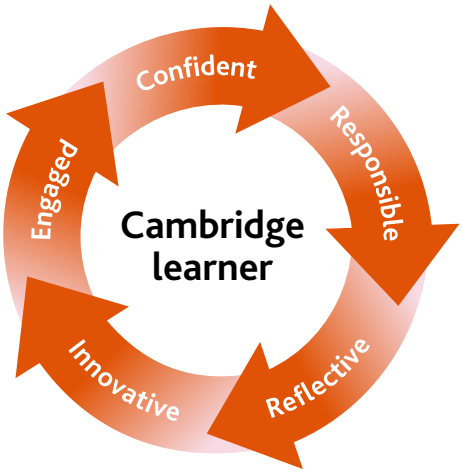
Students can choose from 70 subjects in any combination – it is taught by over 4800 schools in over 150 countries.

Our programmes balance a thorough knowledge and understanding of a subject and help to develop the skills learners need for their next steps in education or employment.

Cambridge IGCSE Combined Science enables learners to:

- increase their understanding of the technological world
- take an informed interest in scientific matters
- recognise the usefulness (and limitations) of scientific method, and how to apply this to other disciplines and in everyday life
- develop relevant attitudes, such as a concern for accuracy and precision, objectivity, integrity, enquiry, initiative and inventiveness
- develop an interest in, and care for, the environment
- better understand the influence and limitations placed on scientific study by society, economy, technology, ethics, the community and the environment
- develop an understanding of the scientific skills essential for both further study and everyday life.

Our approach encourages learners to be:



'The strength of Cambridge IGCSE qualifications is internationally recognised and has provided an international pathway for our students to continue their studies around the world.'

Gary Tan, Head of Schools and CEO, Raffles International Group of Schools, Indonesia

International recognition and acceptance

Our expertise in curriculum, teaching and learning, and assessment is the basis for the recognition of our programmes and qualifications around the world. The combination of knowledge and skills in Cambridge IGCSE Combined Science gives learners a solid foundation for further study. Candidates who achieve grades A* to C are well prepared to follow a wide range of courses including Cambridge International AS & A Level science subjects.

Cambridge IGCSEs are accepted and valued by leading universities and employers around the world as evidence of academic achievement. Many universities require a combination of Cambridge International AS & A Levels and Cambridge IGCSEs or equivalent to meet their entry requirements.

UK NARIC, the national agency in the UK for the recognition and comparison of international qualifications and skills, has carried out an independent benchmarking study of Cambridge IGCSE and found it to be comparable to the standard of the reformed GCSE in the UK. This means students can be confident that their Cambridge IGCSE qualifications are accepted as equivalent to UK GCSEs by leading universities worldwide.

Learn more at **www.cambridgeinternational.org/recognition**

'Cambridge IGCSE is one of the most sought-after and recognised qualifications in the world. It is very popular in Egypt because it provides the perfect preparation for success at advanced level programmes.'

Managing Director of British School in Egypt BSE

Supporting teachers

We provide a wide range of resources, detailed guidance and innovative training and professional development so that you can give your students the best possible preparation for Cambridge IGCSE. To find out which resources are available for each syllabus go to our School Support Hub.

The School Support Hub is our secure online site for Cambridge teachers where you can find the resources you need to deliver our programmes. You can also keep up to date with your subject and the global Cambridge community through our online discussion forums.

Find out more at **www.cambridgeinternational.org/support**

Planning and preparation

- Next step guides
- Schemes of work
- Specimen papers
- Syllabuses
- Teacher guides

Teaching and assessment

- Endorsed resources
- Online forums
- Support for coursework and speaking tests

Support for Cambridge IGCSE

Learning and revision

- Example candidate responses
- Learner guides
- Past papers and mark schemes
- Specimen paper answers

Results

- Candidate Results Service
- Principal examiner reports for teachers
- Results Analysis

Sign up for email notifications about changes to syllabuses, including new and revised products and services at **www.cambridgeinternational.org/syllabusupdates**

Professional development

We support teachers through:

- Introductory Training – face-to-face or online
- Extension Training – face-to-face or online
- Enrichment Professional Development – face-to-face or online

Find out more at **www.cambridgeinternational.org/events**

- Cambridge Professional Development Qualifications

Find out more at **www.cambridgeinternational.org/profdev**

Supporting exams officers

We provide comprehensive support and guidance for all Cambridge exams officers. Find out more at: **www.cambridgeinternational.org/eoguide**

2 Syllabus overview

Aims

The aims describe the purposes of a course based on this syllabus.

You can deliver some of the aims using suitable local, international or historical examples and applications, or through collaborative experimental work.

The aims are to:

- provide an enjoyable and worthwhile educational experience for all learners, whether or not they go on to study science beyond this level
- enable learners to acquire sufficient knowledge and understanding to:
  - become confident citizens in a technological world and develop an informed interest in scientific matters
  - be suitably prepared for studies beyond Cambridge IGCSE
- allow learners to recognise that science is evidence-based and understand the usefulness, and the limitations, of scientific method
- develop skills that:
  - are relevant to the study and practice of science
  - are useful in everyday life
  - encourage a systematic approach to problem-solving
  - encourage efficient and safe practice
  - encourage effective communication through the language of science
- develop attitudes relevant to science such as:
  - concern for accuracy and precision
  - objectivity
  - integrity
  - enquiry
  - initiative
  - inventiveness
- enable learners to appreciate that:
  - science is subject to social, economic, technological, ethical and cultural influences and limitations
  - the applications of science may be both beneficial and detrimental to the individual, the community and the environment.

Cambridge Assessment International Education is an education organisation and politically neutral. The contents of this syllabus, examination papers and associated materials do not endorse any political view. We endeavour to treat all aspects of the exam process neutrally.

Content overview

The subject content is divided into three sections: Biology (B1–B12), Chemistry (C1–C12) and Physics (P1–P6). **Candidates must study all three sections.**

Biology

B1 Characteristics of living organisms
B2 Cells
B3 Biological molecules
B4 Enzymes
B5 Plant nutrition
B6 Animal nutrition
B7 Transport
B8 Gas exchange and respiration
B9 Coordination and response
B10 Reproduction
B11 Organisms and their environment
B12 Human influences on ecosystems

Chemistry

C1 The particulate nature of matter
C2 Experimental techniques
C3 Atoms, elements and compounds
C4 Stoichiometry
C5 Electricity and chemistry
C6 Energy changes in chemical reactions
C7 Chemical reactions
C8 Acids, bases and salts
C9 The Periodic Table
C10 Metals
C11 Air and water
C12 Organic chemistry

Physics

P1 Motion
P2 Work, energy and power
P3 Thermal physics
P4 Properties of waves, including light and sound
P5 Electrical quantities
P6 Electric circuits

Assessment overview

All candidates take three papers.

Candidates who have studied the Core subject content, or who are expected to achieve a grade D or below, should be entered for Paper 1, Paper 3 and either Paper 5 or Paper 6. These candidates will be eligible for grades C to G.

Candidates who have studied the Extended subject content (Core and Supplement), and who are expected to achieve a grade C or above, should be entered for Paper 2, Paper 4 and either Paper 5 or Paper 6. These candidates will be eligible for grades A* to G.

Core candidates take:

Paper 1 45 minutes
Multiple Choice (Core) 30%
40 marks
40 four-option multiple-choice questions
Questions will be based on the Core subject content.
Externally assessed

Extended candidates take:

Paper 2 45 minutes
Multiple Choice (Extended) 30%
40 marks
40 four-option multiple-choice questions
Questions will be based on the Extended subject content (Core and Supplement).
Externally assessed

and Core candidates take:

Paper 3 1 hour 15 minutes
Theory (Core) 50%
80 marks
Short-answer and structured questions
Questions will be based on the Core subject content.
Externally assessed

and Extended candidates take:

Paper 4 1 hour 15 minutes
Theory (Extended) 50%
80 marks
Short-answer and structured questions
Questions will be based on the Extended subject content (Core and Supplement).
Externally assessed

All candidates take either:

Paper 5 1 hour 15 minutes
Practical Test 20%
40 marks
Questions will be based on the experimental skills in section 4.
Externally assessed

or:

Paper 6 1 hour
Alternative to Practical 20%
40 marks
Questions will be based on the experimental skills in section 4.
Externally assessed

Information on availability is in the **Before you start** section.

Assessment objectives

The assessment objectives (AOs) are:

AO1 Knowledge with understanding

Candidates should be able to demonstrate knowledge and understanding of:

- scientific phenomena, facts, laws, definitions, concepts and theories
- scientific vocabulary, terminology and conventions (including symbols, quantities and units)
- scientific instruments and apparatus, including techniques of operation and aspects of safety
- scientific and technological applications with their social, economic and environmental implications.

Subject content defines the factual material that candidates may be required to recall and explain. Candidates will also be asked questions which require them to apply this material to unfamiliar contexts and to apply knowledge from one area of the syllabus to another.

Questions testing this assessment objective will often begin with one of the following words: *define*, *state*, *describe*, *explain* (*using your knowledge and understanding*) or *outline* (see the *Glossary of terms used in science papers*).

AO2 Handling information and problem-solving

Candidates should be able, in words or using other written forms of presentation (i.e. symbolic, graphical and numerical), to:

- locate, select, organise and present information from a variety of sources
- translate information from one form to another
- manipulate numerical and other data
- use information to identify patterns, report trends and draw inferences
- present reasoned explanations for phenomena, patterns and relationships
- make predictions and hypotheses
- solve problems, including some of a quantitative nature.

Questions testing these skills may be based on information that is unfamiliar to candidates, requiring them to apply the principles and concepts from the syllabus to a new situation, in a logical, deductive way.

Questions testing these skills will often begin with one of the following words: *predict*, *suggest*, *calculate* or *determine* (see the *Glossary of terms used in science papers*).

AO3 Experimental skills and investigations

Candidates should be able to:

- demonstrate knowledge of how to safely use techniques, apparatus and materials (including following a sequence of instructions where appropriate)
- plan experiments and investigations
- make and record observations, measurements and estimates
- interpret and evaluate experimental observations and data
- evaluate methods and suggest possible improvements.

Weighting for assessment objectives

The approximate weightings allocated to each of the assessment objectives (AOs) are summarised below.

Assessment objectives as a percentage of the qualification

Assessment objective	Weighting in IGCSE %
AO1 Knowledge with understanding	50
AO2 Handling information and problem-solving	30
AO3 Experimental skills and investigations	20
Total	100

Assessment objectives as a percentage of each component

Assessment objective	Weighting in components %		
	Papers 1 and 2	Papers 3 and 4	Papers 5 and 6
AO1 Knowledge with understanding	63	63	–
AO2 Handling information and problem-solving	37	37	–
AO3 Experimental skills and investigations	–	–	100
Total	100	100	100

3 Subject content

The subject content that follows is divided into three sections: Biology (B1–B12), Chemistry (C1–C12) and Physics (P1–P6). **Candidates must study all three sections.**

All candidates should be taught the Core subject content. Candidates who are only taught the Core subject content can achieve a maximum of grade C. Candidates aiming for grades A* to C should be taught the Extended subject content. The Extended subject content includes both the Core and the Supplement.

This syllabus gives you the flexibility to design a course that will interest, challenge and engage your learners. Where appropriate you are responsible for selecting suitable subject contexts, resources and examples to support your learners' study. These should be appropriate for the learners' age, cultural background and learning context as well as complying with your school policies and local legal requirements.

Scientific subjects are, by their nature, experimental. Learners should pursue a fully integrated course which allows them to develop their practical skills by carrying out practical work and investigations within all of the topics listed.

Biology

B1 Characteristics of living organisms

B1.1 Characteristics of living organisms

Core

1 Describe the characteristics of living organisms by defining the terms:
   - *movement* as an action by an organism causing a change of position or place
   - *respiration* as the chemical reactions in cells that break down nutrient molecules and release energy
   - *sensitivity* as the ability to detect and respond to changes in the environment
   - *growth* as a permanent increase in size
   - *reproduction* as the processes that make more of the same kind of organism
   - *excretion* as removal from organisms of toxic materials and substances in excess of requirements
   - *nutrition* as taking in of materials for energy, growth and development

0653 Biology

B2 Cells

B2.1 Cell structure

Core	Supplement
1 State that living organisms are made of cells	
2 Describe and compare the structure of a plant cell with an animal cell, as seen under a light microscope, limited to cell wall, nucleus, cytoplasm, chloroplasts, vacuoles and location of the cell membrane	
3 State the functions of the structures seen under the light microscope in the plant cell and in the animal cell	4 Relate the structure of the following to their functions: – ciliated cells – movement of mucus in the trachea and bronchi – root hair cells – absorption – palisade mesophyll cells – photosynthesis – red blood cells – transport of oxygen – sperm and egg cells – reproduction
5 Calculate magnification and size of biological specimens using millimetres as units	

B2.2 Movement in and out of cells

Core	Supplement
1 Define *diffusion* as the net movement of particles from a region of their higher concentration to a region of their lower concentration down a concentration gradient, as a result of their random movement	
2 State that substances move into and out of cells by diffusion through the cell membrane	
3 State that water diffuses through partially permeable membranes by osmosis	4 Define *osmosis* as the net movement of water molecules from a region of higher water potential (dilute solution) to a region of lower water potential (concentrated solution), through a partially permeable membrane
5 State that water moves in and out of cells by osmosis through the cell membrane	
6 Investigate and describe the effects on plant tissues of immersing them in solutions of different concentrations	

0653 Biology

B3 Biological molecules

B3.1 Biological molecules

Core

1 List the chemical elements that make up:
   – carbohydrates
   – fats
   – proteins
2 State that large molecules are made from smaller molecules, limited to:
   – starch and glycogen from glucose
   – proteins from amino acids
   – fats and oils from fatty acids and glycerol
3 Describe the use of:
   – iodine solution to test for starch
   – Benedict's solution to test for reducing sugars
   – biuret test for proteins
   – ethanol emulsion test for fats and oils
4 State that water is important as a solvent

B4 Enzymes

B4.1 Enzymes

Core

1 Define *enzymes* as proteins that function as biological catalysts

3 Investigate and describe the effect of changes in temperature and pH on enzyme activity

Supplement

2 Explain enzyme action with reference to the complementary shape of the active site of an enzyme and its substrate and the formation of a product
4 Explain the effect of changes in temperature on enzyme activity, in terms of kinetic energy, shape and fit, frequency of effective collisions and denaturation
5 Explain the effect of changes in pH on enzyme activity in terms of shape and fit and denaturation

0653 Biology

B5 Plant nutrition

B5.1 Plant nutrition

Core	Supplement
1 Define *photosynthesis* as the process by which plants manufacture carbohydrates from raw materials using energy from light	
2 State the word equation for photosynthesis: carbon dioxide + water → glucose + oxygen, in the presence of light and chlorophyll	3 State the balanced equation for photosynthesis $6CO_2 + 6H_2O \xrightarrow[\text{chlorophyll}]{\text{light}} C_6H_{12}O_6 + 6O_2$
	4 Explain that chlorophyll transfers light energy into chemical energy in molecules, for the synthesis of carbohydrates
	5 Outline the subsequent use and storage of the carbohydrates made in photosynthesis
6 Investigate the necessity for chlorophyll, light and carbon dioxide for photosynthesis, using appropriate controls	7 Investigate and describe the effect of varying light intensity and temperature on the rate of photosynthesis (e.g. in submerged aquatic plants)
8 Identify chloroplasts, cuticle, guard cells and stomata, upper and lower epidermis, palisade mesophyll, spongy mesophyll, vascular bundles, xylem and phloem in leaves of a dicotyledonous plant	9 Describe the significance of the features of a leaf in terms of functions, to include: – palisade mesophyll and distribution of chloroplasts – photosynthesis – stomata, spongy mesophyll cells and guard cells – gas exchange – xylem for transport and support – phloem for transport
10 Describe the importance of: – nitrate ions for making amino acids – magnesium ions for making chlorophyll	11 Explain the effects of nitrate ion and magnesium ion deficiency on plant growth

0653 Biology

B6 Animal nutrition

B6.1 Diet

Core

1 State what is meant by the term *balanced diet* for humans

2 List the principal sources of, and describe the dietary importance of:
- carbohydrates
- fats
- proteins
- vitamins, limited to C and D
- mineral salts, limited to calcium and iron
- fibre (roughage)
- water

Supplement

3 Explain how age, gender and activity affect the dietary needs of humans including during pregnancy and while breast-feeding

4 Describe the effects of malnutrition in relation to starvation, constipation, coronary heart disease, obesity and scurvy

5 Explain the causes and effects of vitamin D and iron deficiencies

B6.2 Alimentary canal

Core

1 Define *ingestion* as the taking of substances, e.g. food and drink, into the body through the mouth

2 Define *digestion* as the breakdown of large, insoluble food molecules into small, water-soluble molecules using mechanical and chemical processes

5 Define *absorption* as the movement of small food molecules and ions through the wall of the intestine into the blood

6 Define *egestion* as the passing out of food that has not been digested or absorbed, as faeces, through the anus

7 Identify the main regions of the alimentary canal and associated organs, limited to mouth, salivary glands, oesophagus, stomach, small intestine, pancreas, liver, gall bladder, large intestine and anus

8 Describe the functions of the regions of the alimentary canal listed above, in relation to ingestion, digestion, absorption and egestion of food

Supplement

3 Define *mechanical digestion* as the breakdown of food into smaller pieces without chemical change to the food molecules

4 Define *chemical digestion* as the breakdown of large, insoluble molecules into small, soluble molecules

0653 Biology

B6.3 Digestion

Core	Supplement
1 State the significance of chemical digestion in the alimentary canal in producing small, soluble molecules that can be absorbed	2 State the functions of enzymes as follows: – amylase breaks down starch to simpler sugars – protease breaks down protein to amino acids – lipase breaks down fats to fatty acids and glycerol 3 State where, in the alimentary canal, amylase, protease and lipase are secreted 4 State the functions of the hydrochloric acid in gastric juice, limited to killing bacteria in food and giving an acid pH for enzymes

B7 Transport

B7.1 Transport in plants

Core	Supplement
1 State the functions of xylem and phloem	
2 Identify the position of xylem as seen in sections of roots, stems and leaves, limited to non-woody dicotyledonous plants	
3 Identify root hair cells, as seen under the light microscope, and state their functions	4 Explain that the large surface area of root hairs increases the rate of the absorption of water
5 State the pathway taken by water through root, stem and leaf as root hair cell, root cortex cells, xylem and mesophyll cells	
6 Investigate, using a suitable stain, the pathway of water through the above-ground parts of a plant	
7 State that water is transported from the roots to leaves through the xylem vessels	
8 Define *transpiration* as loss of water vapour from plant leaves by evaporation of water at the surfaces of the mesophyll cells followed by diffusion of water vapour through the stomata	
9 Investigate and describe the effects of variation of temperature and humidity on transpiration rate	10 Explain the effects of variation of temperature and humidity on transpiration rate

0653 Biology

B7.2 Transport in mammals

Core

1 Describe the circulatory system as a system of blood vessels with a pump and valves to ensure one-way flow of blood

4 Name and identify the structures of the mammalian heart, limited to the muscular wall, the septum, the left and right ventricles and atria, one-way valves and coronary arteries

5 State that blood is pumped away from the heart into arteries and returns to the heart in veins

7 Name the main blood vessels to and from the:
- heart, limited to vena cava, aorta, pulmonary artery and pulmonary vein
- lungs, limited to the pulmonary artery and pulmonary vein

9 Investigate and state the effect of physical activity on pulse rate

11 Describe the structure and functions of arteries, veins and capillaries

13 List the components of blood as red blood cells, white blood cells, platelets and plasma

14 Identify red and white blood cells, as seen under the light microscope, on prepared slides and in diagrams and photomicrographs

15 State the functions of the following components of blood:
- red blood cells in transporting oxygen, including the role of haemoglobin
- white blood cells in phagocytosis and antibody production
- platelets in clotting (details are **not** required)
- plasma in the transport of blood cells, ions, soluble nutrients, hormones and carbon dioxide

Supplement

2 Describe double circulation in terms of circulation to the lungs and circulation to the body tissues in mammals

3 Explain the advantages of a double circulation

6 Describe the functioning of the heart in terms of the contraction of muscles of the atria and ventricles and the action of the valves

8 Describe coronary heart disease in terms of the blockage of coronary arteries and state the possible risk factors as diet, stress, smoking, genetic predisposition, age and gender

10 Explain the effect of physical activity on the heart rate

12 Explain how the structures of arteries, veins and capillaries are adapted for their function

0653 Biology

B8 Gas exchange and respiration

B8.1 Gas exchange

Core

1 Name and identify the lungs, diaphragm, ribs, intercostal muscles, larynx, trachea, bronchi, bronchioles, alveoli and associated capillaries

3 State the differences in composition between inspired and expired air limited to oxygen, carbon dioxide and water vapour

5 Use limewater as a test for carbon dioxide to investigate the differences in composition between inspired and expired air

6 Investigate and describe the effects of physical activity on rate and depth of breathing

Supplement

2 List the features of gas exchange surfaces in animals, limited to large surface area, thin surface, good blood supply and good ventilation with air

4 Explain the differences in composition between inspired and expired air

7 Explain the effects of physical activity on rate and depth of breathing in terms of the increased carbon dioxide concentration in the blood, causing an increased rate of breathing

8 Explain the role of goblet cells, mucus and ciliated cells in protecting the gas exchange system from pathogens and particles

9 State that tobacco smoking can cause chronic obstructive pulmonary disease (COPD), lung cancer and coronary heart disease

10 Describe the effects on the gas exchange system of tobacco smoke and its major toxic components, limited to carbon monoxide, nicotine and tar

B8.2 Respiration

Core

1 State the uses of energy in the body of humans limited to: muscle contraction, protein synthesis, growth and the maintenance of a constant body temperature

3 State the word equation for aerobic respiration as glucose + oxygen → carbon dioxide + water

Supplement

2 Define *aerobic respiration* as the chemical reactions in cells that use oxygen to break down nutrient molecules to release energy

4 State the balanced chemical equation for aerobic respiration as

$$C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$$

0653 Biology

B9 Coordination and response

B9.1 Hormones in humans

Core

1 Define a *hormone* as a chemical substance, produced by a gland, carried by the blood, which alters the activity of one or more specific target organs

2 Describe adrenaline as the hormone secreted in 'fight or flight' situations and its effects, limited to increased breathing and pulse rate and widened pupils

4 Give examples of situations in which adrenaline secretion increases

Supplement

3 Discuss the role of the hormone adrenaline in the chemical control of metabolic activity, including increasing the blood glucose concentration and pulse rate

B9.2 Tropic responses

Core

1 Define *gravitropism* as a response in which parts of a plant grow towards or away from gravity

2 Define *phototropism* as a response in which parts of a plant grow towards or away from the direction from which light is coming

4 Investigate gravitropism and phototropism in shoots and roots

Supplement

3 Explain phototropism and gravitropism of a shoot as examples of the chemical control of plant growth

5 Explain the role of auxin in controlling shoot growth, limited to:

- auxin made in shoot tip (only)
- auxin spreads through the plant from the shoot tip
- auxin is unequally distributed in response to light and gravity
- auxin stimulates cell elongation

B10 Reproduction

B10.1 Asexual and sexual reproduction

Core

1 Define *asexual reproduction* as a process resulting in the production of genetically identical offspring from one parent

2 Identify examples of asexual reproduction from information provided

3 Define *sexual reproduction* as a process involving the fusion of the nuclei of two gametes (sex cells) to form a zygote and the production of offspring that are genetically different from each other

0653 Biology

B10.2 Sexual reproduction in plants

Core

1 Identify and draw, using a hand lens if necessary, the sepals, petals, stamens, filaments and anthers, carpels, style, stigma, ovary and ovules, of an insect-pollinated flower

3 State the functions of the sepals, petals, anthers, stigmas and ovaries

5 Define *pollination* as the transfer of pollen grains from the anther to the stigma

6 Name the agents of pollination

7 State that fertilisation occurs when a pollen nucleus fuses with a nucleus in an ovule

9 Investigate and state the environmental conditions that affect germination of seeds: limited to the requirement for water, oxygen and a suitable temperature

Supplement

2 Use a hand lens to identify and describe the anthers and stigmas of a wind-pollinated flower

4 Distinguish between the pollen grains of insect-pollinated and wind-pollinated flowers

8 Describe the structural adaptations of insect-pollinated and wind-pollinated flowers

B10.3 Sexual reproduction in humans

Core

1 Identify and name on diagrams of the male reproductive system: the testes, scrotum, sperm ducts, prostate gland, urethra and penis

2 State the function of the parts of the male reproductive system limited to:

- testes – production of male gametes (sperm)
- scrotum – sac that holds the testes outside the body
- sperm ducts – transfer sperm to the urethra
- prostate gland – secrete fluids for sperm to swim in forming semen
- urethra – carries urine and semen out of the body
- penis – transfers semen to vagina during sexual intercourse

3 Identify and name on diagrams of the female reproductive system: the ovaries, oviducts, uterus, cervix and vagina

continued

0653 Biology

B10.3 Sexual reproduction in humans continued

Core	Supplement
4 State the function of the parts of the female reproductive system limited to: – ovaries – release of female gametes (eggs) – oviducts – transfers egg to uterus and the site of fertilisation – uterus – where the fetus develops – cervix – ring of muscle at the opening of the uterus – vagina – receives penis during sexual intercourse	
5 Describe fertilisation as the fusion of the nuclei from a male gamete (sperm) and a female gamete (egg cell/ovum)	6 Compare male and female gametes in terms of size, structure, motility and numbers 7 State the adaptive features of sperm, limited to flagellum and the presence of enzymes 8 State the adaptive features of egg cells, limited to energy stores and a jelly coating that changes after fertilisation
9 Describe the menstrual cycle in terms of changes in the ovaries and in the lining of the uterus (knowledge of sex hormones is **not** required)	
10 State that in early development, the zygote forms an embryo which is a ball of cells that implants into the wall of the uterus	11 State the functions of the umbilical cord, placenta, amniotic sac and amniotic fluid 12 Describe the function of the placenta and umbilical cord in relation to exchange of dissolved nutrients, gases and excretory products and providing a barrier to toxins (structural details are **not** required)
13 State that human immunodeficiency virus (HIV) infection may lead to acquired immune deficiency syndrome (AIDS) 14 Describe the methods of transmission of HIV 15 Explain how the spread of sexually transmitted infections (STIs) is controlled	

0653 Biology

B11 Organisms and their environment

B11.1 Organisms and their environment

Core	Supplement
1 State that the Sun is the principal source of energy input to biological systems	
2 Define the terms: – *food chain* as showing the transfer of energy from one organism to the next, beginning with a producer – *food web* as a network of interconnected food chains – *producer* as an organism that makes its own organic nutrients, usually using energy from sunlight, through photosynthesis – *consumer* as an organism that gets its energy by feeding on other organisms – *herbivore* as an animal that gets its energy by eating plants – *carnivore* as an animal that gets its energy by eating other animals – *decomposer* as an organism that gets its energy from dead or waste organic matter	3 Define the terms: – *ecosystem* as a unit containing all of the organisms and their environment, interacting together, in a given area, e.g. a lake – *trophic level* as the position of an organism in a food chain or food web
	4 Describe how energy is transferred between trophic levels
	5 Explain why food chains usually have fewer than five trophic levels
6 Construct simple food chains	
7 Interpret food chains and food webs in terms of identifying producers and consumers	
8 State that consumers may be classed as primary, secondary and tertiary according to their position in a food chain	9 Identify producers, primary consumers, secondary consumers, tertiary consumers and quaternary consumers as the trophic levels in food webs and food chains

0653 Biology

B12 Human influences on ecosystems

B12.1 Human influences on ecosystems

Core

1 Describe the carbon cycle, limited to photosynthesis, respiration, feeding, decomposition, fossilisation and combustion

3 List the undesirable effects of deforestation as an example of habitat destruction, to include extinction, loss of soil, flooding and increase of carbon dioxide in the atmosphere

Supplement

2 Discuss the effects of the combustion of fossil fuels and the cutting down of forests on the oxygen and carbon dioxide concentrations in the atmosphere

4 Explain the process of eutrophication of water in terms of:

- increased availability of nitrate and other ions
- increased growth of producers
- increased decomposition after death of producers
- increased aerobic respiration by decomposers
- reduction in dissolved oxygen
- death of organisms requiring dissolved oxygen in water

Chemistry

C1 The particulate nature of matter

C1.1 The particulate nature of matter

Core

1 State the distinguishing properties of solids, liquids and gases

2 Describe the structure of solids, liquids and gases in terms of particle separation, arrangement and types of motion

3 Describe the changes of state in terms of melting, boiling, evaporation, freezing and condensation

5 Describe qualitatively the pressure and temperature of a gas in terms of the motion of its particles

6 Demonstrate understanding of the terms *atom*, *molecule* and *ion*

Supplement

4 Explain changes of state in terms of particle theory and the energy changes involved

C2 Experimental techniques

C2.1 Measurement

Core

1 Name and suggest appropriate apparatus for the measurement of time, temperature, mass and volume, including burettes, pipettes and measuring cylinders

C2.2 Criteria of purity

Core

1 Interpret simple chromatograms

Supplement

2 Interpret simple chromatograms, including the use of R_f values

C2.3 Methods of purification

Core

1 Describe and explain methods of separation and purification by the use of a suitable solvent, filtration, crystallisation, distillation, fractional distillation and paper chromatography

2 Suggest suitable separation and purification techniques, given information about the substances involved

C3 Atoms, elements and compounds

C3.1 Physical and chemical changes

Core

1 Identify physical and chemical changes, and understand the differences between them

C3.2 Elements, compounds and mixtures

Core

1 Describe the differences between elements, mixtures and compounds, and between metals and non-metals

2 Define the terms *solvent*, *solute*, *solution* and *concentration*

0653 Chemistry

C3.3 Atomic structure and the Periodic Table

Core

1 Describe the structure of an atom in terms of a central nucleus, containing protons and neutrons, and 'shells' of electrons

2 Describe the build-up of electrons in 'shells' and understand the significance of the noble gas electronic structures and of the outer shell electrons

(The ideas of the distribution of electrons in s and p orbitals and in d-block elements are **not** required)

3 State the charges and approximate relative masses of protons, neutrons and electrons

4 Define and use *proton number* (atomic number) as the number of protons in the nucleus of an atom

5 Define and use *nucleon number* (mass number) as the total number of protons and neutrons in the nucleus of an atom

Note: a copy of the Periodic Table, as shown in the Appendix, will be provided in Papers 1, 2, 3 and 4.

Supplement

6 Use proton number and the simple structure of atoms to explain the basis of the Periodic Table, with special reference to the elements of proton numbers 1 to 20

C3.4 Ions and ionic bonds

Core	Supplement
1 Describe the formation of ions by electron loss or gain	
2 Use dot-and-cross diagrams to describe the formation of ionic bonds between Group I and Group VII	3 Describe the formation of ionic bonds between metallic and non-metallic elements to include the strong attraction between ions because of their opposite electrical charges
	4 Describe the lattice structure of ionic compounds as a regular arrangement of alternating positive and negative ions, exemplified by the sodium chloride structure

C3.5 Molecules and covalent bonds

Core	Supplement
1 State that non-metallic elements form simple molecules with covalent bonds between atoms	
2 Describe the formation of single covalent bonds in H_2, Cl_2, H_2O, CH_4, NH_3 and HCl as the sharing of pairs of electrons leading to the noble gas configuration including the use of dot-and-cross diagrams	3 Use and draw dot-and-cross diagrams to represent the bonding in the more complex covalent molecules such as N_2, C_2H_4, CH_3OH, and CO_2
4 Describe the differences in volatility, solubility and electrical conductivity between ionic and covalent compounds	5 Explain the differences in melting point and boiling point of ionic and covalent compounds in terms of attractive forces

C4 Stoichiometry

C4.1 Stoichiometry

Core	Supplement
1 Use the symbols of the elements and write the formulae of simple compounds	2 Determine the formula of an ionic compound from the charges on the ions present
3 Deduce the formula of a simple compound from the relative numbers of atoms present	
4 Deduce the formula of a simple compound from a model or a diagrammatic representation	
5 Construct and use word equations	
6 Interpret and balance simple symbol equations	7 Construct and use symbol equations, with state symbols, including ionic equations

C5 Electricity and chemistry

C5.1 Electricity and chemistry

Core	Supplement
1 Define *electrolysis* as the breakdown of an ionic compound when molten or in aqueous solution by the passage of electricity	
2 Use the terms *inert electrode*, *electrolyte*, *anode* and *cathode*	3 Describe electrolysis in terms of the ions present and the reactions at the electrodes, in terms of gain of electrons by cations and loss of electrons by anions to form atoms
4 Describe the electrode products and the observations made, using inert electrodes (platinum or carbon), in the electrolysis of: – molten lead(II) bromide – concentrated aqueous sodium chloride – dilute sulfuric acid	5 Predict the products of the electrolysis of a specified molten binary compound

0653 Chemistry

C6 Energy changes in chemical reactions

C6.1 Energy changes in chemical reactions

Core	Supplement
1 Describe the meaning of *exothermic* and *endothermic* reactions	2 Describe bond breaking as an endothermic process and bond forming as an exothermic process
	3 Draw and label energy level diagrams for exothermic and endothermic reactions using data provided
	4 Interpret energy level diagrams showing exothermic and endothermic reactions and the activation energy of a reaction

C7 Chemical reactions

C7.1 Rate (speed) of reaction

Core

1 Describe practical methods for investigating the rate of a reaction which produces a gas

2 Interpret data obtained from experiments concerned with rate of reaction

4 Describe the effect of concentration, particle size, catalysts and temperature on the rate of reactions

Note: Candidates should be encouraged to use the term *rate* rather than *speed*.

Supplement

3 Suggest suitable apparatus, given information, for experiments, including collection of gases and measurement of rates of reaction

5 Describe and explain the effect of changing concentration in terms of frequency of collisions between reacting particles

6 Describe and explain the effect of changing temperature in terms of the frequency of collisions between reacting particles and more colliding particles possessing the minimum energy (activation energy) to react

C7.2 Redox

Core

1 Describe oxidation and reduction in chemical reactions in terms of oxygen loss/gain (Oxidation state limited to its use to name ions, e.g. iron(II), iron(III), copper(II).)

Supplement

2 Define and identify an *oxidising agent* as a substance which oxidises another substance during a redox reaction and a *reducing agent* as a substance which reduces another substance during a redox reaction

0653 Chemistry

C8 Acids, bases and salts

C8.1 The characteristic properties of acids and bases

Core

1 Describe neutrality and relative acidity and alkalinity in terms of pH (whole numbers only) measured using universal indicator

2 Describe the characteristic properties of acids (exemplified by dilute hydrochloric acid and dilute sulfuric acid) including their effect on litmus paper and their reactions with metals, bases and carbonates

3 Describe and explain the importance of controlling acidity in soil

C8.2 Preparation of salts

Core

1 Describe the preparation, separation and purification of salts using techniques specified in Section C2 and the reactions specified in Section C8.1

Supplement

2 Suggest a method of making a given salt from suitable starting material, given appropriate information

C8.3 Identification of ions and gases

Core

1 Describe and use the following tests to identify:

aqueous cations:

ammonium, calcium, copper(II), iron(II), iron(III) and zinc, by means of aqueous sodium hydroxide and aqueous ammonia as appropriate (formulae of complex ions are **not** required).

cations:

flame tests to identify lithium, sodium, potassium and copper(II)

anions:

carbonate (by reaction with dilute acid and then limewater), chloride (by reaction under acidic conditions with aqueous silver nitrate), nitrate (by reduction with aluminium) and sulfate (by reaction under acidic conditions with aqueous barium ions)

gases:

ammonia (using damp red litmus paper), carbon dioxide (using limewater), chlorine (using damp litmus paper), hydrogen (using a lighted splint), oxygen (using a glowing splint)

0653 Chemistry

C9 The Periodic Table

C9.1 The Periodic Table

Core

1 Describe the Periodic Table as a method of classifying elements and its use to predict properties of elements

C9.2 Periodic trends

Core

1 Describe the change from metallic to non-metallic character across a period

Supplement

2 Describe and explain the relationship between group number, number of outer-shell electrons and metallic/non-metallic character

C9.3 Group properties

Core

1 Describe lithium, sodium and potassium in Group I as a collection of relatively soft metals showing a trend in melting point, density and reaction with water

3 Describe the halogens, chlorine, bromine and iodine in Group VII, as a collection of diatomic non-metals showing a trend in colour and physical state

Supplement

2 Predict the properties of other elements in Group I, given data, where appropriate

4 State the reaction of chlorine, bromine and iodine with other halide ions

5 Predict the properties of other elements in Group VII, given data where appropriate

6 Identify trends in other groups, given data about the elements concerned

C9.4 Transition elements

Core

1 Describe the transition elements as a collection of metals having high densities, high melting points and forming coloured compounds, and which, as elements and compounds, often act as catalysts

C9.5 Noble gases

Core

1 Describe the noble gases, in Group VIII or 0, as being unreactive, monoatomic gases and explain this in terms of electronic structure

2 State the uses of the noble gases in providing an inert atmosphere, i.e. argon in lamps, helium for filling balloons

0653 Chemistry

C10 Metals

C10.1 Properties of metals

Core

1 Describe the general physical properties of metals as solids with high melting and boiling points, malleable and good conductors of heat and electricity

2 Describe alloys, such as brass, as mixtures of a metal with other elements

3 Explain in terms of their properties why alloys are used instead of pure metals

Supplement

4 Identify representations of alloys from diagrams of structure

C10.2 Reactivity series

Core

1 Place in order of reactivity: potassium, sodium, calcium, magnesium, aluminium, (carbon), zinc, iron, (hydrogen) and copper, by reference to the reactions, if any, of the elements with:
   - water or steam
   - dilute hydrochloric acid
   - reduction of their oxides with carbon

3 Deduce an order of reactivity from a given set of experimental results

Supplement

2 Describe the reactivity series in terms of the tendency of a metal to form its positive ion, illustrated by its reaction, if any, with the aqueous ions of other listed metals

C10.3 Extraction of metals from their ores

Core

1 Describe the use of carbon in the extraction of copper from copper oxide

3 Know that aluminium is extracted from the ore bauxite by electrolysis

5 Describe metal ores as a finite resource and hence the need to recycle metals

Supplement

2 Describe and explain the essential reactions in the extraction of iron from hematite in the blast furnace

$C + O_2 \rightarrow CO_2$

$C + CO_2 \rightarrow 2CO$

$Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$

4 Relate the method of extraction of a metal from its ore to its position in the reactivity series for the metals listed in Section C10.2 and for other metals, given information

0653 Chemistry

C11 Air and water

C11.1 Water

Core

1 Describe a chemical test for water using copper(II) sulfate and cobalt(II) chloride

2 Describe, in outline, the treatment of the water supply in terms of filtration and chlorination

C11.2 Air

Core

1 State the composition of clean air as being a mixture of 78% nitrogen, 21% oxygen and small quantities of noble gases, water vapour and carbon dioxide

2 Name the common pollutants in air as being carbon monoxide, sulfur dioxide and oxides of nitrogen

3 State the adverse effect of these common air pollutants on buildings and on health

4 State the conditions required for the rusting of iron (presence of oxygen and water)

5 Describe and explain barrier methods of rust prevention, including paint and other coatings

C11.3 Carbon dioxide and methane

Core

1 State the formation of carbon dioxide:
- as a product of complete combustion of carbon-containing substances
- as a product of respiration
- as a product of the reaction between an acid and a carbonate
- as a product of thermal decomposition of calcium carbonate

2 State that carbon dioxide and methane are greenhouse gases

Supplement

3 State that increased concentrations of greenhouse gases cause an enhanced greenhouse effect, which may contribute to climate change

0653 Chemistry

C12 Organic chemistry

C12.1 Fuels

Core

1 State that coal, natural gas and petroleum are fossil fuels that produce carbon dioxide on combustion

2 Name methane as the main constituent of natural gas

3 Describe petroleum as a mixture of hydrocarbons and its separation into useful fractions by fractional distillation

5 Name the uses of the fractions as:

- refinery gas for bottled gas for heating and cooking
- gasoline fraction for fuel (petrol) in cars
- naphtha fraction as a feedstock for making chemicals
- diesel oil / gas oil for fuel in diesel engines
- bitumen for road surfaces

Supplement

4 Describe the properties of molecules within a fraction

C12.2 Homologous series

Supplement

1 Describe the homologous series of alkanes and alkenes as families of compounds with the same general formula and similar chemical properties

C12.3 Alkanes

Core

1 Describe alkanes as saturated hydrocarbons whose molecules contain only single covalent bonds

2 Describe the properties of alkanes (exemplified by methane) as being generally unreactive, except in terms of burning

3 Describe the complete combustion of hydrocarbons to give carbon dioxide and water

0653 Chemistry

C12.4 Alkenes

Core

1 Describe alkenes as unsaturated hydrocarbons whose molecules contain one double covalent bond

2 State that cracking is a reaction that produces alkenes

4 Recognise saturated and unsaturated hydrocarbons:
- from molecular structures
- by their reaction with aqueous bromine

5 Describe the formation of poly(ethene) as an example of addition polymerisation of monomer units

Supplement

3 Describe the formation of smaller alkanes, alkenes and hydrogen by the cracking of larger alkane molecules and state the conditions required for cracking

Physics

P1 Motion

P1.1 Length and time

Core

1 Use and describe the use of rules and measuring cylinders to find a length or a volume
2 Use and describe the use of clocks and devices, both analogue and digital, for measuring an interval of time
3 Obtain an average value for a small distance and for a short interval of time by measuring multiples (including the period of a pendulum)

P1.2 Motion

Core

1 Define *speed* and calculate average speed from $\frac{\text{total distance}}{\text{total time}}$
2 Plot and interpret a speed–time graph and a distance-time graph
3 Recognise from the shape of a speed–time graph when a body is:
   – at rest
   – moving with constant speed
   – moving with changing speed
8 Demonstrate understanding that acceleration and deceleration are related to changing speed including qualitative analysis of the gradient of a speed–time graph

Supplement

4 Calculate the area under a speed–time graph to work out the distance travelled for motion with constant acceleration
5 Calculate acceleration from the gradient of a speed–time graph
6 Recognise linear motion for which the acceleration is constant and calculate the acceleration
7 Recognise motion for which the acceleration is not constant

P1.3 Mass and weight

Core

1 Distinguish between *mass* and *weight*

2 Know that the Earth is the source of a gravitational field

4 Recognise that g is the gravitational force on unit mass and is measured in N / kg

5 Recall and use the equation $W = mg$

Supplement

3 Describe, and use the concept of, weight as the effect of a gravitational field on a mass

P1.4 Density

Core

1 Recall and use the equation $\rho = \frac{m}{V}$

2 Describe an experiment to determine the density of a liquid and of a regularly shaped solid and make the necessary calculation

Supplement

3 Describe the determination of the density of an irregularly shaped solid by the method of displacement and make the necessary calculation

0653 Physics

P1.5 Effects of forces

Core

1 Describe how forces may change the size, shape and motion of a body

5 Understand friction as the force between two surfaces which impedes motion and results in heating

6 Recognise air resistance as a form of friction

7 Find the resultant of two or more forces acting along the same line

8 Recognise that if there is no resultant force on a body it either remains at rest or continues at constant speed in a straight line

Supplement

2 Plot and interpret extension–load graphs and describe the associated experimental procedure

3 State Hooke's law and recall and use the expression $F = kx$, where k is the spring constant

4 Recognise the significance of the term 'limit of proportionality' for an extension–load graph

P1.6 Pressure

Core

1 Relate qualitatively pressure to force and area, using appropriate examples

Supplement

2 Recall and use the equation $p = F / A$

P2 Work, energy and power

P2.1 Work

Core	Supplement
1 Relate (without calculation) work done to the magnitude of a force and distance moved in the direction of the force	2 Recall and use $W = Fd = \Delta E$

0653 Physics

P2.2 Energy

Core	Supplement
1 Demonstrate an understanding that work done = energy transferred	
2 Demonstrate understanding that an object may have energy due to its motion (kinetic energy, KE) or its position (potential energy, PE) and that energy may be transferred and stored	
3 Give and identify examples of changes in kinetic, gravitational potential, chemical potential, elastic potential (strain), thermal, sound and electrical potential energy that have occurred as a result of an event or process	4 Recall and use the expressions KE = $\frac{1}{2}mv^2$ and gravitational potential energy (GPE) = mgh or change in GPE = $mg\Delta h$
5 Recognise that energy is transferred during events and processes, including examples of transfer by forces (mechanical working), by electric currents (electrical working), by heating and by waves	
6 Apply the principle of conservation of energy to simple examples	

P2.3 Power

Core	Supplement
1 Relate (without calculation) power to work done and time taken, using appropriate examples	2 Recall and use the equation $P = \Delta E / t$ in simple systems, including electrical circuits

P2.4 Energy resources

Core	Supplement
1 Distinguish between renewable and non-renewable sources of energy	
2 Describe how electricity or other useful forms of energy may be obtained from: – chemical energy stored in fuel – water, including the energy stored in waves, in tides, and in water behind hydroelectric dams – geothermal resources – nuclear fission – heat and light from the Sun (solar cells and panels) – wind energy	3 Give advantages and disadvantages of each method in terms of renewability, cost, reliability, scale and environmental impact 4 Understand that the Sun is the source of energy for all our energy resources except geothermal, nuclear and tidal 5 Understand that the source of tidal energy is mainly the moon 6 Show an understanding that energy is released by nuclear fusion in the Sun

0653 Physics

P3 Thermal physics

P3.1 Simple kinetic molecular model of matter

Core	Supplement
1 State the distinguishing properties of solids, liquids and gases	2 Relate the properties of solids, liquids and gases to the forces and distances between the molecules and to the motion of the molecules
3 Describe qualitatively the molecular structure of solids, liquids and gases in terms of the arrangement, separation, and motion of the molecules	
4 Describe qualitatively the pressure of a gas and the temperature of a gas, liquid or solid in terms of the motion of its particles	
5 Use and describe the use of thermometers to measure temperature on the Celsius scale	
6 State the meaning of *melting point* and *boiling point*, and recall the melting and boiling points for water	
7 Describe evaporation in terms of the escape of more-energetic molecules from the surface of a liquid	
8 Relate evaporation to the consequent cooling of the liquid	9 Demonstrate an understanding of how temperature, surface area and draught over a surface influence evaporation

P3.2 Matter and thermal properties

Core

1 Describe qualitatively the thermal expansion of solids, liquids and gases at constant pressure

2 Identify and explain some of the everyday applications and consequences of thermal expansion

P3.3 Thermal processes

P3.3.1 Conduction

Core

1 Recognise and name typical good and bad thermal conductors

2 Describe experiments to demonstrate the properties of good and bad thermal conductors

Supplement

3 Explain conduction in solids in terms of molecular vibrations and transfer by electrons

P3.3.2 Convection

Core

1 Recognise convection as the main method of energy transfer in fluids

3 Interpret and describe experiments designed to illustrate convection in liquids and gases (fluids)

Supplement

2 Relate convection in fluids to density changes

P3.3.3 Radiation

Core

1 Recognise radiation as the method of energy transfer that does not require a medium to travel through

2 Identify infrared radiation as the part of the electromagnetic spectrum often involved in energy transfer by radiation

Supplement

3 Describe the effect of surface colour (black or white) and texture (dull or shiny) on the emission, absorption and reflection of radiation

4 Interpret and describe experiments to investigate the properties of good and bad emitters and good and bad absorbers of infrared radiation

P3.3.4 Consequences of energy transfer

Core

1 Identify and explain some of the everyday applications and consequences of conduction, convection and radiation

0653 Physics

P4 Properties of waves, including light and sound

P4.1 General wave properties

Core

1 Demonstrate understanding that waves transfer energy without transferring matter

2 Describe what is meant by *wave motion* as illustrated by vibration in ropes and springs and by experiments using water waves

3 State the meaning of *speed*, *frequency*, *wavelength* and *amplitude*

5 Describe how waves can undergo:
- reflection at a plane surface
- refraction due to a change of speed

Supplement

4 Distinguish between transverse and longitudinal waves and give suitable examples

6 Recall and use the equation $v = f\lambda$

7 Understand that refraction is caused by a change in speed as a wave moves from one medium to another

P4.2 Light

P4.2.1 Reflection of light

Core

1 Describe the formation of an optical image by a plane mirror and give its characteristics

2 Recall and use the law

angle of incidence i = angle of reflection r

recognising these angles are measured to the normal

3 Perform simple constructions, measurements and calculations for reflection by plane mirrors

P4.2.2 Refraction of light

Core

1 Interpret and describe an experimental demonstration of the refraction of light

P4.2.3 Thin converging lens

Core

1 Describe the action of a thin converging lens on a beam of light

2 Use the terms *principal focus* and *focal length*

3 Draw ray diagrams for the formation of a real image by a single lens

Supplement

4 Use and describe the use of a single lens as a magnifying glass

0653 Physics

P4.3 Electromagnetic spectrum

Core

1 Describe the main features of the electromagnetic spectrum in order of frequency, from radio waves to gamma radiation (γ)

2 State that all electromagnetic waves travel with the same high speed in a vacuum and approximately the same in air

4 Describe typical properties and uses of radiations in all the different regions of the electromagnetic spectrum including:
- radio and television communications (radio waves)
- satellite television and telephones (microwaves)
- electrical appliances, remote controllers for televisions and intruder alarms (infrared)
- medicine and security (X-rays)

5 Demonstrate an understanding of safety issues regarding the use of microwaves and X-rays

6 State the dangers of ultraviolet radiation, from the Sun or from tanning lamps

Supplement

3 State that the speed of electromagnetic waves in a vacuum is 3.0×10^8 m / s

P4.4 Sound

Core

1 Describe the production of sound by vibrating sources

4 State that the approximate range of audible frequencies for a healthy human ear is 20 Hz to 20 000 Hz

5 Show an understanding that a medium is needed to transmit sound waves

6 Describe and interpret an experiment to determine the speed of sound in air, including calculation

8 Relate the loudness and pitch of sound waves to amplitude and frequency

9 Describe how the reflection of sound may produce an echo

Supplement

2 Describe the longitudinal nature of sound waves

3 Describe the transmission of sound waves in air in terms of compressions and rarefactions

7 Recognise that sound travels faster in liquids than in gases and faster in solids than in liquids

0653 Physics

P5 Electrical quantities

P5.1 Electric charge

Core

1 State that there are positive and negative charges
2 State that unlike charges attract and that like charges repel
3 Describe and interpret simple experiments to show the production and detection of electrostatic charges by friction
4 State that charging a body involves the addition or removal of electrons
5 Distinguish between electrical conductors and insulators and give typical examples

P5.2 Current, potential difference and electromotive force (e.m.f)

Core

1 Demonstrate understanding of *current*, *potential difference*, *e.m.f.* and *resistance*
2 State that current is related to the flow of charge
5 State that current in metals is due to a flow of electrons
6 State that the potential difference (p.d.) across a circuit component is measured in volts
7 Use and describe the use of an ammeter and a voltmeter, both analogue and digital
8 State that the electromotive force (e.m.f.) of an electrical source of energy is measured in volts

Supplement

3 Know and use the formula $Q = It$
4 Show understanding that a current is a rate of flow of charge and recall and use the equation $I = Q / t$

P5.3 Resistance

Core

1 State that resistance = p.d. / current and understand qualitatively how changes in p.d. or resistance affect current
2 Recall and use the equation $R = V / I$

Supplement

3 Recall and use quantitatively the proportionality between resistance and length, and the inverse proportionality between resistance and cross-sectional area of a wire

0653 Physics

P6 Electric circuits

P6.1 Circuit diagrams

Core

1 Draw and interpret circuit diagrams containing sources, switches, resistors (fixed and variable), lamps, ammeters, voltmeters and fuses

(Symbols for other common circuit components will be provided in questions)

P6.2 Series and parallel circuits

Core

1 Understand that the current at every point in a series circuit is the same

2 Calculate the combined resistance of two or more resistors in series

4 State that, for a parallel circuit, the current from the source is larger than the current in each branch

6 State that the combined resistance of two resistors in parallel is less than that of either resistor by itself

8 State the advantages of connecting lamps in parallel in a circuit

Supplement

3 Recall and use the fact that the sum of the p.d.s across the components in a series circuit is equal to the total p.d. across the supply

5 Recall and use the fact that the current from the source is the sum of the currents in the separate branches of a parallel circuit

7 Calculate the combined resistance of two resistors in parallel

P6.3 Electrical energy

Supplement

1 Recall and use the equations $P = IV$ and $E = IVt$

P6.4 Dangers of electricity

Core

1 Identify electrical hazards including:
- damaged insulation
- overheating of cables
- damp conditions

2 State that a fuse protects a circuit

3 Explain the use of fuses and choose appropriate fuse ratings

4 Details of the assessment

All candidates take three papers.

Candidates who have studied the Core subject content, or who are expected to achieve a grade D or below, should be entered for Paper 1, Paper 3 and either Paper 5 or Paper 6. These candidates will be eligible for grades C to G.

Candidates who have studied the Extended subject content (Core and Supplement), and who are expected to achieve a grade C or above, should be entered for Paper 2, Paper 4 and either Paper 5 or Paper 6. These candidates will be eligible for grades A* to G.

Core assessment

Core candidates take the following papers that have questions based on the Core subject content only:

Paper 1 Multiple Choice (Core)

45 minutes, 40 marks

Forty compulsory multiple-choice items of the four-option type. This paper tests assessment objectives AO1 and AO2.

Paper 3 Theory (Core)

1 hour 15 minutes, 80 marks

Short-answer and structured questions testing assessment objectives AO1 and AO2.

Extended assessment

Extended candidates take the following papers that have questions based on the Core and Supplement subject content:

Paper 2 Multiple Choice (Extended)

45 minutes, 40 marks

Forty compulsory multiple-choice items of the four-option type. This paper tests assessment objectives AO1 and AO2.

Paper 4 Theory (Extended)

1 hour 15 minutes, 80 marks

Short-answer and structured questions testing assessment objectives AO1 and AO2.

Practical assessment

All candidates take one practical component from a choice of two:

Paper 5 Practical Test

1 hour 15 minutes, 40 marks

This paper tests assessment objective AO3 in a practical context.

or

Paper 6 Alternative to Practical

1 hour, 40 marks

This paper tests assessment objective AO3 in a written paper.

Whichever practical paper you choose, please be aware that:

- they test the same assessment objective, AO3
- they require the same experimental skills to be learned and developed
- the same sequence of practical activities is appropriate.

Candidates must not use textbooks or any of their course notes in the practical component.

These papers are based on testing experimental skills. One question on each paper assesses the skill of planning. This question will be based on any one of the sciences, which could be: Biology, Chemistry or Physics.

Questions in the practical papers are structured to assess performance across the full grade range. The information candidates need to answer the questions is in the question paper itself or the experimental context and skills listed below. The questions do not assess specific subject content.

Experimental skills tested in Paper 5 Practical Test and Paper 6 Alternative to Practical

Candidates may be required to do the following:

- carefully follow a sequence of instructions
- describe, explain or comment on experimental arrangements and techniques
- select the most appropriate apparatus or method for a task and justify the choice made
- draw, complete or label diagrams of apparatus
- perform simple arithmetical calculations
- take readings from an appropriate measuring device or from an image of the device (e.g. thermometer, rule, protractor, measuring cylinder, ammeter, stop-watch), including:
  - reading analogue and digital scales with accuracy and appropriate precision
  - interpolating between scale divisions when appropriate
  - correcting for zero errors when appropriate
- plan to take a sufficient number and range of measurements, repeating where appropriate to obtain an average value
- describe or explain precautions taken in carrying out a procedure to ensure safety or the accuracy of observations and data, including the control of variables and repetition of measurements

- identify key variables and describe how, or explain why, certain variables should be controlled
- record observations systematically, for example in a table, using appropriate units and to a consistent and appropriate degree of precision
- process data, using a calculator where necessary
- present and analyse data graphically, including the use of best-fit lines where appropriate, interpolation and extrapolation, and the determination of a gradient, intercept or intersection
- interpret and evaluate observations and experimental data
- draw an appropriate conclusion, justifying it by reference to the data and using an appropriate explanation
- comment critically on a procedure or point of practical detail, and suggest an appropriate improvement
- evaluate the quality of data, identifying and dealing appropriately with any anomalous results
- identify possible causes of uncertainty, in data or in a conclusion
- make estimates or describe outcomes which demonstrate their familiarity with an experiment, procedure or technique
- plan an experiment or investigation, including making reasoned predictions of expected results and suggesting suitable apparatus and techniques.

Biology

Candidates may be asked questions on the following experimental contexts:

- the use of familiar, and unfamiliar, techniques to record observations and data, process them and make deductions from them
- recall of simple chemical tests, e.g. for food substances and the use of limewater, hydrogencarbonate indicator, litmus and universal indicator paper
- recognise, observe, record and measure images of familiar, and unfamiliar, biological specimens
- make a clear line drawing from an image of a specimen, calculating the magnification and adding labels as required.

Chemistry

Candidates may be asked questions on the following experimental contexts:

- simple quantitative experiments involving the measurement of volumes and/or masses
- rates (speeds) of reaction
- measurement of temperature based on a thermometer with 1 °C graduations and energetics
- problems of an investigatory nature, possibly including suitable organic compounds
- filtration
- electrolysis
- identification of ions and gases
- metals and the reactivity series
- acids, bases, oxides and preparation of salts
- redox reactions and rusting.

Physics

Candidates may be asked questions on the following experimental contexts:

- measurement of physical quantities such as length or volume or force or density
- cooling and heating
- springs and balances
- timing motion or oscillations
- electrical circuits, circuit diagrams and electrical symbols
- optics equipment such as mirrors, prisms and lenses
- procedures using simple apparatus, in situations where the method may not be familiar to the candidate
- use or describe the use of common techniques, apparatus and materials, e.g. ray-tracing equipment or the connection of electric circuits
- explain the manipulation of the apparatus to obtain observations or measurements, e.g.:
  - when determining a derived quantity, such as the extension per unit load for a spring
  - when testing/identifying the relationship between two variables, such as between the p.d. across a wire and its length
  - when comparing physical quantities, such as two masses, using a balancing method.

Teaching experimental skills

We expect you to look for suitable opportunities to embed practical techniques and investigative work throughout the course.

The best way to prepare candidates for these papers is to integrate practical work fully into the course so that it becomes a normal part of your teaching. Practical work helps candidates to:

- develop a deeper understanding of the syllabus topics
- learn to appreciate the way in which scientific theories are developed and tested
- develop experimental skills and positive scientific attitudes such as objectivity, integrity, cooperation, enquiry and inventiveness.

Note on taking readings

When approximate volumes are used, e.g. about 2 cm^3, it is expected that candidates will estimate this and not use measuring devices.

A measuring instrument should be used to its full precision. Thermometers may be marked in 1 °C intervals but it is often appropriate to interpolate between scale divisions and record a temperature to the nearest 0.0 °C or 0.5 °C. Measurements using a rule require suitable accuracy of recording, such as 15.0 cm rather than 15 cm; the use of millimetres when appropriate should be encouraged. Similarly, when measuring current, it is often more appropriate to use milliamperes rather than amperes.

Apparatus list

This list contains the items you are likely to need for teaching the experimental skills needed for both practical papers, as well as the Paper 5 exam. It is not exhaustive and does not include equipment commonly regarded as standard in a science laboratory. The *Confidential Instructions* we send you before the Paper 5 exam will give the detailed requirements for that exam.

- rulers capable of measuring to 1 mm
- metre rule
- mounted needles or seekers or long pins with large heads
- means of cutting biological materials, such as scalpels, solid edged razor blades or knives
- scissors
- forceps
- means of writing on glassware
- beakers, 100 cm^3, 250 cm^3
- polystyrene or other plastic beakers of approximate capacity 150 cm^3
- test-tubes (Pyrex or hard glass), approximately 125 mm × 16 mm
- boiling tubes, approximately 150 mm × 25 mm
- delivery tubes
- conical flasks, within the range 150 cm^3 to 250 cm^3
- means of measuring small volumes of liquids, such as syringes (with needles removed)
- measuring cylinders, 100 cm^3, 50 cm^3, 25 cm^3, 10 cm^3
- dropping pipettes
- white tiles
- spotting tiles
- water-bath
- large containers (e.g. plastic bowl) to hold cold water
- hand lens ×6 magnification
- thermometers, –10 °C to +110 °C with 1 °C graduations
- stop-clocks (or wall-clock or wrist-watch), to measure to an accuracy of 1 s
- Petri dishes
- glass rods
- spatulas
- wooden splints
- chemicals (e.g. for food tests, limewater test)
- indicators (e.g. litmus paper, universal indicator paper, full range universal indicator, hydrogencarbonate indicator)
- burettes, 50 cm^3
- pipettes, 25 cm^3
- pipette fillers
- filter funnels and filter paper
- wash bottle
- ammeter FSD 1 A, 1.5 A
- voltmeter FSD 1 V, 5 V
- electrical cells (batteries) and holders to enable several cells to be joined

- connecting leads and crocodile clips
- d.c. power supply, variable to 12 V
- low-voltage filament lamps in holders
- various resistors and resistance wire
- switches
- good supply of masses and holders
- 2 cm expendable springs
- clamps and stands
- pendulum bobs
- newton meters
- Plasticine or modelling clay
- wooden boards
- converging lens with $f = 15$ cm
- glass or Perspex block, rectangular and semi-circular
- glass or Perspex prism, triangular
- optics pins
- plane mirrors
- ray box

Glossary of terms used in science papers

This glossary (which is relevant only to science subjects) will prove helpful to candidates as a guide, but it is neither exhaustive nor definitive. The glossary has been deliberately kept brief, not only with respect to the number of terms included, but also to the descriptions of their meanings. Candidates should appreciate that the meaning of a term must depend, in part, on its context.

1 *Define* (the term(s) …) is intended literally, only a formal statement or equivalent paraphrase being required.

2 *What do you understand by/What is meant by* (the term(s) …) normally implies that a definition should be given, together with some relevant comment on the significance or context of the term(s) concerned, especially where two or more terms are included in the question. The amount of supplementary comment intended should be interpreted in the light of the indicated mark value.

3 *State* implies a concise answer with little or no supporting argument (e.g. a numerical answer that can readily be obtained 'by inspection').

4 *List* requires a number of points, generally each of one word, with no elaboration. Where a given number of points is specified this should not be exceeded.

5 (a) *Explain* may imply reasoning or some reference to theory, depending on the context. It is another way of asking candidates to give reasons. The candidate needs to leave the examiner in no doubt why something happens.

(b) *Give a reason / Give reasons* is another way of asking candidates to explain why something happens.

6 *Describe* requires the candidate to state in words (using diagrams where appropriate) the main points.

Describe and *explain* may be coupled, as may *state* and *explain*.

7 *Discuss* requires the candidate to give a critical account of the points involved.

8 *Outline* implies brevity (i.e. restricting the answer to giving essentials).

9 *Predict* implies that the candidate is expected to make a prediction not by recall but by making a logical connection between other pieces of information.

10 *Deduce* implies that the candidate is not expected to produce the required answer by recall but by making a logical connection between other pieces of information.

11 *Suggest* is used in two main contexts, i.e. either to imply that there is no unique answer (e.g. in Physics there are several examples of energy resources from which electricity, or other useful forms of energy, may be obtained), or to imply that candidates are expected to apply their general knowledge of the subject to a 'novel' situation, one that may be formally 'not in the syllabus' – many data response and problem-solving questions are of this type.

12 *Find* is a general term that may variously be interpreted as *calculate*, *measure*, *determine*, etc.

13 *Calculate* is used when a numerical answer is required. In general, working should be shown, especially where two or more steps are involved.

14 *Measure* implies that the quantity concerned can be directly obtained from a suitable measuring instrument (e.g. length using a rule, or mass using a balance).

15 *Determine* often implies that the quantity concerned cannot be measured directly but is obtained from a graph or by calculation.

16 *Estimate* implies a reasoned order of magnitude statement or calculation of the quantity concerned, making such simplifying assumptions as may be necessary about points of principle and about the values of quantities not otherwise included in the question.

17 *Sketch*, when applied to graph work, implies that the shape and/or position of the curve need only be qualitatively correct, but candidates should be aware that, depending on the context, some quantitative aspects may be looked for (e.g. passing through the origin, having an intercept).

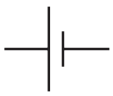
In diagrams, *sketch* implies that simple, free-hand drawing is acceptable; nevertheless, care should be taken over proportions and the clear exposition of important details.

5 Appendix

Electrical symbols

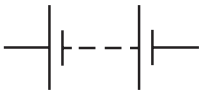
Candidates are expected to be able to recall and use the standard electrical symbols listed below.

cell



battery of cells

or

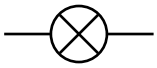


power supply

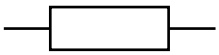
a.c. power supply

junction of conductors

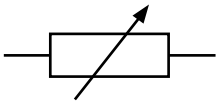
lamp



fixed resistor

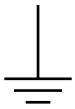


variable resistor



switch

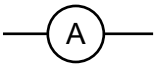
earth or ground



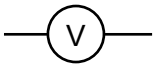
electric bell

motor

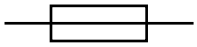
ammeter



voltmeter



fuse



heater

Symbols and units for physical quantities

Candidates should be able to give the symbols for the following physical quantities and, where indicated, state the units in which they are measured. The list for the Extended subject content includes both the Core and the Supplement.

Candidates should be familiar with the following multipliers: M mega, k kilo, c centi, m milli.

Core			Supplement		
Quantity	***Usual symbol***	***Usual unit***	***Quantity***	***Usual symbol***	***Usual unit***
length	*l, h* …	km, m, cm, mm			
area	*A*	m^2, cm^2			
volume	*V*	m^3, cm^3			
weight	*W*	N			
mass	*m, M*	kg, g	mass	*m, M*	mg
time	*t*	h, min, s	time	*t*	ms
density	ρ	g/cm^3, kg/m^3			
speed	*u, v*	km/h, m/s, cm/s			
acceleration	*a*		acceleration	*a*	m/s^2
acceleration of free fall	*g*		acceleration of free fall	*g*	m/s^2
force	*F*	N			
gravitational field strength	*g*	N/kg			
work done	*W, E*	J, kJ, MJ			
energy	*E*	J, kJ, MJ			
power	*P*	W, kW, MW			
pressure	*p*	N/m^2	pressure	*p*	Pa
temperature	θ, *T*	°C			
frequency	*f*	Hz, kHz			
wavelength	λ	m, cm			
focal length	*f*	cm			
angle of incidence	*i*	degree (°)			
angle of reflection, refraction	*r*	degree (°)			
potential difference/voltage	*V*	V, mV			
current	*I*	A, mA			
e.m.f.	*E*	V			
resistance	*R*	Ω			
			charge	*Q*	C

Notes for use in qualitative analysis

Tests for anions

anion	test	test result
carbonate (CO_3^{2-})	add dilute acid	effervescence, carbon dioxide produced
chloride (Cl^-) [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	white ppt.
nitrate (NO_3^-) [in solution]	add aqueous sodium hydroxide, then aluminium foil; warm carefully	ammonia produced
sulfate (SO_4^{2-}) [in solution]	acidify, then add aqueous barium nitrate	white ppt.

Tests for aqueous cations

cation	effect of aqueous sodium hydroxide	effect of aqueous ammonia
ammonium (NH_4^+)	ammonia produced on warming	–
calcium (Ca^{2+})	white ppt., insoluble in excess	no ppt. or very slight white ppt.
copper(II) (Cu^{2+})	light blue ppt., insoluble in excess	light blue ppt., soluble in excess, giving a dark blue solution
iron(II) (Fe^{2+})	green ppt., insoluble in excess	green ppt., insoluble in excess
iron(III) (Fe^{3+})	red-brown ppt., insoluble in excess	red-brown ppt., insoluble in excess
zinc (Zn^{2+})	white ppt., soluble in excess, giving a colourless solution	white ppt., soluble in excess, giving a colourless solution

Tests for gases

gas	test and test result
ammonia (NH_3)	turns damp red litmus paper blue
carbon dioxide (CO_2)	turns limewater milky
chlorine (Cl_2)	bleaches damp litmus paper
hydrogen (H_2)	'pops' with a lighted splint
oxygen (O_2)	relights a glowing splint

Flame tests for metal ions

metal ion	flame colour
lithium (Li^+)	red
sodium (Na^+)	yellow
potassium (K^+)	lilac
copper(II) (Cu^{2+})	blue-green

The Periodic Table

Group																	
I	II											III	IV	V	VI	VII	VIII
							1 H hydrogen 1										2 He helium 4
3 Li lithium 7	4 Be beryllium 9											5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	10 Ne neon 20
11 Na sodium 23	12 Mg magnesium 24											13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium –	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131
55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids	72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium –	85 At astatine –	86 Rn radon –
87 Fr francium –	88 Ra radium –	89–103 actinoids	104 Rf rutherfordium –	105 Db dubnium –	106 Sg seaborgium –	107 Bh bohrium –	108 Hs hassium –	109 Mt meitnerium –	110 Ds darmstadtium –	111 Rg roentgenium –	112 Cn copernicium –		114 Fl flerovium –		116 Lv livermorium –		

Key

atomic number
atomic symbol
name
relative atomic mass

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium –	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
actinoids	89 Ac actinium –	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium –	94 Pu plutonium –	95 Am americium –	96 Cm curium –	97 Bk berkelium –	98 Cf californium –	99 Es einsteinium –	100 Fm fermium –	101 Md mendelevium –	102 No nobelium –	103 Lr lawrencium –

The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.)

Back to contents page

Safety in the laboratory

Responsibility for safety matters rests with centres. Further information can be found from the following UK associations, publications and regulations.

Associations

CLEAPSS is an advisory service providing support in practical science and technology.
www.cleapss.org.uk

Publications

CLEAPSS Laboratory Handbook, updated 2009 (available to CLEAPSS members only)
CLEAPSS Hazcards, 2007 update of 1995 edition (available to CLEAPSS members only)

UK regulations

Control of Substances Hazardous to Health Regulations (COSHH) 2002 and subsequent amendment in 2004
www.legislation.gov.uk/uksi/2002/2677/contents/made
www.legislation.gov.uk/uksi/2004/3386/contents/made

A brief guide may be found at
www.hse.gov.uk/pubns/indg136.pdf

Mathematical requirements

Calculators may be used in all parts of the examination.

Candidates should be able to:

- add, subtract, multiply and divide
- use averages, decimals, fractions, percentages, ratios and reciprocals
- use standard notation, including both positive and negative indices
- understand significant figures and use them appropriately
- recognise and use direct and inverse proportion
- use positive, whole number indices in algebraic expressions
- draw charts and graphs from given data
- interpret charts and graphs
- determine the gradient and intercept of a graph
- select suitable scales and axes for graphs
- make approximate evaluations of numerical expressions
- recall and use equations for the areas of a rectangle, triangle and circle and the volumes of a rectangular block and a cylinder
- use mathematical instruments (ruler, compasses, protractor and set square)
- understand the meaning of angle, curve, circle, radius, diameter, circumference, square, parallelogram, rectangle and diagonal
- solve equations of the form $x = y + z$ and $x = yz$ for any one term when the other two are known
- recognise and use clockwise and anticlockwise directions
- recognise and use points of the compass (N, S, E, W)
- use sines and inverse sines (Extended candidates only).

Presentation of data

The solidus (/) is to be used for separating the quantity and the unit in tables, graphs and charts, e.g. time / s for time in seconds.

(a) Tables

- Each column of a table should be headed with the physical quantity and the appropriate unit, e.g. time / s.
- The column headings of the table can then be directly transferred to the axes of a constructed graph.

(b) Graphs

- Unless instructed otherwise, the independent variable should be plotted on the *x*-axis (horizontal axis) and the dependent variable plotted on the *y*-axis (vertical axis).
- Each axis should be labelled with the physical quantity and the appropriate unit, e.g. time / s.
- The scales for the axes should allow more than half of the graph grid to be used in both directions, and be based on sensible ratios, e.g. 2 cm on the graph grid representing 1, 2 or 5 units of the variable.
- The graph is the whole diagrammatic presentation, including the best-fit line when appropriate. It may have one or more sets of data plotted on it.
- Points on the graph should be clearly marked as crosses (×) or encircled dots (⊙).
- Large 'dots' are penalised. Each data point should be plotted to an accuracy of better than one half of each of the smallest squares on the grid.
- A best-fit line (trend line) should be a single, thin, smooth straight-line or curve. The line does not need to coincide exactly with any of the points; where there is scatter evident in the data, examiners would expect a roughly even distribution of points either side of the line over its entire length. Points that are clearly anomalous should be ignored when drawing the best-fit line.
- The gradient of a straight line should be taken using a triangle whose hypotenuse extends over at least half of the length of the best-fit line, and this triangle should be marked on the graph.

(c) Numerical results

- Data should be recorded so as to reflect the precision of the measuring instrument.
- The number of significant figures given for calculated quantities should be appropriate to the least number of significant figures in the raw data used.

(d) Pie charts

- These should be drawn with the sectors in rank order, largest first, beginning at 'noon' and proceeding clockwise. Pie charts should preferably contain no more than six sectors.

(e) Bar charts

- These should be drawn when one of the variables is not numerical. They should be made up of narrow blocks of equal width that do **not** touch.

(f) Histograms

- These should be drawn when plotting frequency graphs with continuous data. The blocks should be drawn in order of increasing or decreasing magnitude and they **should** touch.

ICT opportunities

In order to play a full part in modern society, candidates need to be confident and effective users of ICT. This syllabus provides candidates with a wide range of opportunities to use ICT in their study of biology, chemistry and physics.

Opportunities for ICT include:

- gathering information from the internet, DVDs and CD-ROMs
- gathering data using sensors linked to data-loggers or directly to computers
- using spreadsheets and other software to process data
- using animations and simulations to visualise scientific ideas
- using software to present ideas and information on paper and on screen.

Conventions (e.g. signs, symbols, terminology and nomenclature)

Syllabuses and question papers conform with generally accepted international practice. In particular, the following document, produced by the Association for Science Education (ASE), should be used as a guideline.

- *Signs, Symbols and Systematics: The ASE Companion to 16–19 Science (2000)*

Litre/dm^3

To avoid any confusion concerning the symbol for litre, **dm^3** will be used in place of *l* or litre.

Decimal markers

In accordance with current ASE convention, decimal markers in examination papers will be a single dot on the line. Candidates are expected to follow this convention in their answers.

Numbers

Numbers from 1000 to 9999 will be printed without commas or spaces. Numbers greater than or equal to 10 000 will be printed without commas. A space will be left between each group of three whole numbers, e.g. 4 256 789.

6 What else you need to know

This section is an overview of other information you need to know about this syllabus. It will help to share the administrative information with your exams officer so they know when you will need their support. Find more information about our administrative processes at **www.cambridgeinternational.org/eoguide**

Before you start

Previous study

We recommend that learners starting this course should have studied a science curriculum such as the Cambridge Lower Secondary programme or equivalent national educational framework.

Guided learning hours

We design Cambridge IGCSE syllabuses based on learners having about 130 guided learning hours for each subject during the course but this is for guidance only. The number of hours a learner needs to achieve the qualification may vary according to local practice and their previous experience of the subject.

Availability and timetables

All Cambridge schools are allocated to one of six administrative zones. Each zone has a specific timetable.

You can view the timetable for your administrative zone at **www.cambridgeinternational.org/timetables**

You can enter candidates in the June and November exam series. If your school is in India, you can also enter your candidates in the March exam series.

Check you are using the syllabus for the year the candidate is taking the exam.

Private candidates can enter for this syllabus. For more information, please refer to the *Cambridge Guide to Making Entries*.

Combining with other syllabuses

Candidates can take this syllabus alongside other Cambridge International syllabuses in a single exam series. The only exceptions are:

- Cambridge IGCSE Biology (0610)
- Cambridge IGCSE (9–1) Biology (0970)
- Cambridge IGCSE Chemistry (0620)
- Cambridge IGCSE (9–1) Chemistry (0971)
- Cambridge IGCSE Physics (0625)
- Cambridge IGCSE (9–1) Physics (0972)
- Cambridge IGCSE Co-ordinated Sciences (Double Award) (0654)
- Cambridge IGCSE (9–1) Co-ordinated Sciences (Double Award) (0973)
- Cambridge IGCSE Physical Science (0652)
- Cambridge O Level Physics (5054)
- Cambridge O Level Chemistry (5070)
- Cambridge O Level Biology (5090)
- Cambridge O Level Combined Science (5129)
- syllabuses with the same title at the same level.

Cambridge IGCSE, Cambridge IGCSE (9–1) and Cambridge O Level syllabuses are at the same level.

Group awards: Cambridge ICE

Cambridge ICE (International Certificate of Education) is a group award for Cambridge IGCSE. It allows schools to offer a broad and balanced curriculum by recognising the achievements of learners who pass examinations in a range of different subjects.

Learn more about Cambridge ICE at **www.cambridgeinternational.org/cambridgeice**

Making entries

Exams officers are responsible for submitting entries to Cambridge International. We encourage them to work closely with you to make sure they enter the right number of candidates for the right combination of syllabus components. Entry option codes and instructions for submitting entries are in the *Cambridge Guide to Making Entries*. Your exams officer has a copy of this guide.

Exam administration

To keep our exams secure, we produce question papers for different areas of the world, known as administrative zones. We allocate all Cambridge schools to one administrative zone determined by their location. Each zone has a specific timetable. Some of our syllabuses offer candidates different assessment options. An entry option code is used to identify the components the candidate will take relevant to the administrative zone and the available assessment options.

Support for exams officers

We know how important exams officers are to the successful running of exams. We provide them with the support they need to make your entries on time. Your exams officer will find this support, and guidance for all other phases of the Cambridge Exams Cycle, at **www.cambridgeinternational.org/eoguide**

Retakes

Candidates can retake the whole qualification as many times as they want to. This is a linear qualification so candidates cannot re-sit individual components.

Equality and inclusion

In our effort to comply with the UK Equality Act (2010) we have taken all reasonable steps to avoid any direct and indirect discrimination.

The standard assessment arrangements may present barriers for candidates with impairments. Where a candidate is eligible, we may be able to make arrangements to enable that candidate to access assessments and receive recognition of their attainment. We do not agree access arrangements if they give candidates an unfair advantage over others or if they compromise the standards being assessed.

Candidates who cannot access the assessment of any component may be able to receive an award based on the parts of the assessment they have completed.

Information on access arrangements is in the *Cambridge Handbook* at **www.cambridgeinternational.org/eoguide**

Language

This syllabus and the related assessment materials are available in English only.

After the exam

Grading and reporting

Grades A*, A, B, C, D, E, F or G indicate the standard a candidate achieved at Cambridge IGCSE.

A* is the highest and G is the lowest. 'Ungraded' means that the candidate's performance did not meet the standard required for grade G. 'Ungraded' is reported on the statement of results but not on the certificate.

In specific circumstances your candidates may see one of the following letters on their statement of results:

- Q (PENDING)
- X (NO RESULT).

These letters do not appear on the certificate.

On the statement of results and certificates, Cambridge IGCSE is shown as INTERNATIONAL GENERAL CERTIFICATE OF SECONDARY EDUCATION (IGCSE).

How students and teachers can use the grades

Assessment at Cambridge IGCSE has two purposes:

- to measure learning and achievement

  The assessment:

  – confirms achievement and performance in relation to the knowledge, understanding and skills specified in the syllabus.
- to show likely future success

  The outcomes:

  – help predict which students are well prepared for a particular course or career and/or which students are more likely to be successful

  – help students choose the most suitable course or career.

Changes to this syllabus for 2023 and 2024

The syllabus has been updated. This is version 1, published September 2020.

There are no significant changes which affect teaching.

You must read the whole syllabus before planning your teaching programme.

Any textbooks endorsed to support the syllabus for examination from 2019 are still suitable for use with this syllabus.

‘While studying Cambridge IGCSE and Cambridge International A Levels, students broaden their horizons through a global perspective and develop a lasting passion for learning.’

Zhai Xiaoning, Deputy Principal, The High School Affiliated to Renmin University of China

Cambridge Assessment International Education
The Triangle Building, Shaftesbury Road, Cambridge, CB2 8EA, United Kingdom
Tel: +44 (0)1223 553554 Fax: +44 (0)1223 553558
Email: info@cambridgeinternational.org www.cambridgeinternational.org